

ARSENIC REMOVAL FROM WATER USING A NEW CLASS OF MATERIALS WITH ADSORBENT PROPERTIES

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ABSTRACT

One of the strategies for sustainable development is to promote a quality health care system, available to all without discrimination, and improving protection against health threats. In this context, arsenic removal from groundwater for drinking purposes presents challenges at national and global levels. Thus, the present article focuses on removing arsenic from groundwaters by using a new class of materials based on cellulose modified with crown ether (dibenzo-18-crown-6) doped with iron ions. Using such extractants involves only a small amount of crown ether, indicating higher efficiency of produced material, and in order to improve the adsorbent properties and selectivity for arsenic removal, the modified cellulose was functionalized with iron ions.

The new adsorbent material was characterized by using energy-dispersive X-ray analysis and Fourier-transform infrared spectroscopy. To investigate its adsorption properties for arsenic removal, equilibrium, kinetic and thermodynamic studies were performed. Arsenic adsorption from water onto new class of adsorbent material was studied under different experimental conditions such as reaction time, initial arsenic concentration and temperature. Kinetic of adsorption process was better described by pseudo-second-order model. The equilibrium adsorption data were well described by the Sips adsorption isotherm. The values of thermodynamic parameters (ΔG° , ΔH° , ΔS°) showed that the adsorption process was endothermic and spontaneous.

The possibility of reuse of the adsorbent material through adsorption and desorption cycles was also studied, and it was found that the material can be used in three adsorption–desorption cycles.

Keywords: crown ethers, arsenic, adsorption, iron ions, water.

1 INTRODUCTION

Arsenic removal from water is an important and urgent preoccupation and, in this sense, nowadays a series of techniques for arsenic removal are studied and developed. Arsenic is the 20th element and it is found in groundwaters, which became a major problem for the world due to its toxicity especially for countries like Bangladesh, India, China and Japan [1,2]. Bangladesh and provinces adjacent to India such as Bengal are the most affected regions of the world, in which the drinking water of more than 100 million people was contaminated with arsenic and 21% of deaths were reported because of arsenic poisoning [3–6]. Many industries release arsenic without adequate treatment of the water sources. In various regions, the anthropic source was registered as a major source for a higher concentration of arsenic in groundwater [1], and the arsenic content in drinking water over maximum admitted level by the World Health Organization ($10 \mu\text{g As L}^{-1}$) has a negative impact on human health [4–8]. For this reason, arsenic removal from water is an important and urgent preoccupation, and in this way a series of techniques for removing arsenic have been developed. These include precipitation–filtration processes [9]; coagulation–precipitation processes [10], chemical oxidation, oxidation–coagulation or oxidation–precipitation [11]; electrocoagulation [12]; membrane separation such as nanofiltration, reverse osmosis, electrodialysis [13–15], ion exchange [16] and adsorption [17–22].

Among arsenic removal procedures, adsorption is a process with a good efficiency, but an important role is played by the adsorbent material used into the process. Thus, it is very important to develop new methods to produce adsorbent materials with weight adsorption capacity by chemical modification of inorganic and organic solid supports, through functionalization with different pendant groups. Natural polymers as adsorbent materials for arsenic removal from water present a special attention [23–25]. From environmental point of view, the most common and most cost-effective biopolymer available from natural sources is cellulose [26]. Cellulose has proved to be an excellent example in terms of versatility for surface modification [27–29]. Modifying the cellulose structure with crown ethers might be a new method successfully used for metal ion removal from water due to crown ether properties which allow the complexation and embedding of metallic ions inside their structure or at their surface [30]. Researches performed during the last decade proved that the iron ions increase the affinity of adsorbent materials towards arsenic ions.

Based on that, the goals of this article were to obtain a new material with improved adsorbent properties by functionalization cellulose using crown ethers as extractants, followed by loading with iron. The article consists in extended studies of our previous work [31] regarding the chemical modification of cellulose through functionalization with crown ethers and loaded with iron ions for arsenic adsorption from waters. The physical–chemical modification of cellulose by functionalization with dibenzo-18-crown-6 and loaded with Fe(III) ions represent a good choice for As(V) removal from aqueous solutions, due to the fact that cellulose is a relatively inexpensive support and the amount of crown ether necessary for the functionalization is very small, but with representative effect for the adsorption process.

2 MATERIALS AND METHODS

In order to obtain the new class of adsorbent material by chemical modification of solid support, the following were used: as support – cellulose (AVICEL 101, microcrystalline, powder, Sigma-Aldrich, Merck) with a particle size of ~ 50 μm , as extractant – dibenzo-18-crown-6 (DB18C6) (purity, 98%, Sigma-Aldrich, Merck, Germany) and iron chloride as iron sources purchased from Sigma-Aldrich.

2.1 Functionalization of polymer

To obtain the adsorbent material, 0.05 g of dibenzo-18-crown-6 (DB18C6) was accurately weighed and dissolved in 25 mL nitrobenzene (Sigma-Aldrich, Merck, analytical standard). The obtained extractant solution is added over 5 g of support cellulose with a particle size of ~50 μm . The Solvent Impregnated Resins (SIR) dry method was used to functionalize the support. The functionalization is presented schematically in Fig. 1. Thus, the extractant and the support remain in contact for 24 h, after which it is filtered and dried in oven at 50°C for 24 h.

To load the material with iron ions 25 mL of FeCl_3 solution with concentration of 100 mg L^{-1} was added, left in contact for 24 hours, then filtered and dried in the oven for 24 h.

2.2. Characterization of the functionalized polymer

The obtained material was characterized by X-ray energy-dispersive spectroscopy (EDX) using a FEI Quanta FEG 250 Scanning Electron Microscope and Fourier-transform infrared spectroscopy (FTIR) using a Bruker Platinum ATR-QL Diamond instrument in the range of 4,000–400 cm^{-1} .

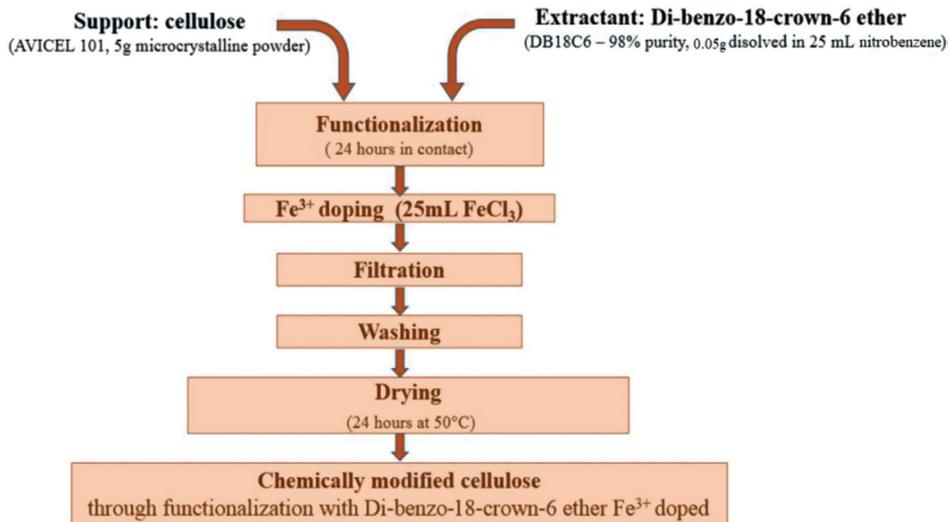


Figure 1: The principle steps for functionalization of the solid support.

2.3 Sorption studies

In order to determine the optimum conditions for the arsenic adsorption process, experiments regarding contact time, initial concentration of As(V) and temperature were carried out. Therefore, a Julabo SW23 thermostatic and shaking water bath with a stirring speed of 200 rpm was used.

To establish the influence of contact time and temperature on the adsorption capacity of the adsorbent material, 0.1 g of material was mixed with 25 mL As(V) solution with concentration of $50 \mu\text{g L}^{-1}$. Samples were shaken for different time periods (0.5, 1, 2, 3, 4, 5, 6, 7 and 8 h) in a thermostatic bath at different temperatures (298K, 308K and 318K) with a stirring speed of 200 rpm. Afterwards, all samples were filtered and the filtrate was analysed in order to evaluate the arsenic residual concentration using an inductively coupled plasma mass spectrometer ICP-MS Bruker Aurora M90. In order to establish the effect of the initial concentration of As(V) on the adsorption capacity of materials, different concentrations of As(V) solutions (10, 25, 50, 75, 100, 150, 175 and $200 \mu\text{g L}^{-1}$) were prepared. Adsorption processes were carried out at pH ~ 7 , for a contact time of 4 h at 298 K.

3 RESULTS AND DISCUSSION

3.1. Characterization of the materials

3.1.1 X-ray energy-dispersive spectroscopy

Cellulose functionalized with DB18C6 and loaded with iron ions was characterized by using X-ray EDX technique. The obtained data are presented in Fig. 2.

From the EDX spectra and quantification, the presence of specific peaks ascribed to elements found in the chemical structure of cellulose and crown ether, namely C and O, might be observed as well as iron.

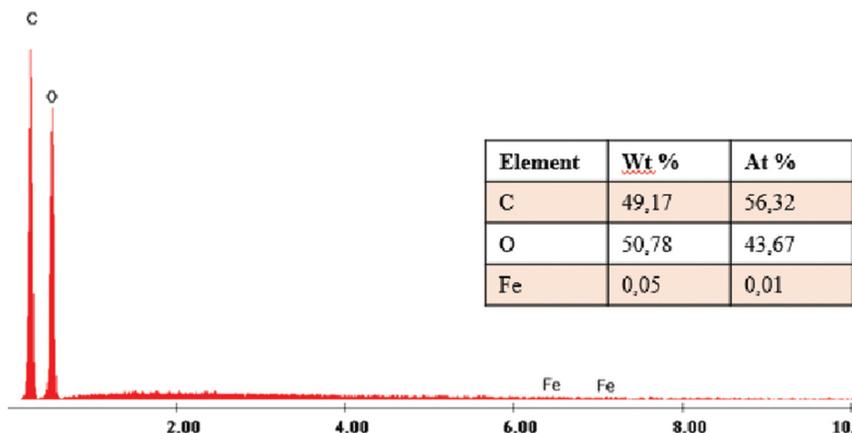


Figure 2: EDX spectra of materials obtained by functionalized cellulose with crown ether DB18C6 and iron ions.

3.1.2 Fourier-transform infrared spectroscopy

FTIR spectroscopy is used to identify chemical groups as well for quantitative analysis of the new material. The FTIR spectra for cellulose functionalized with DB18C6 and iron are shown in Fig. 3.

From FTIR spectra, bands specific for cellulose groups were identified in the range of 3,200–3,400 cm^{-1} associated with the stretching vibrations of O-H bonds [30]. A band located around 1,600 cm^{-1} was as well identified, and it is characteristic for stretching vibrations of water and specific O-H bonds. The bands located at 1,720 and 1,600 cm^{-1} are assigned to

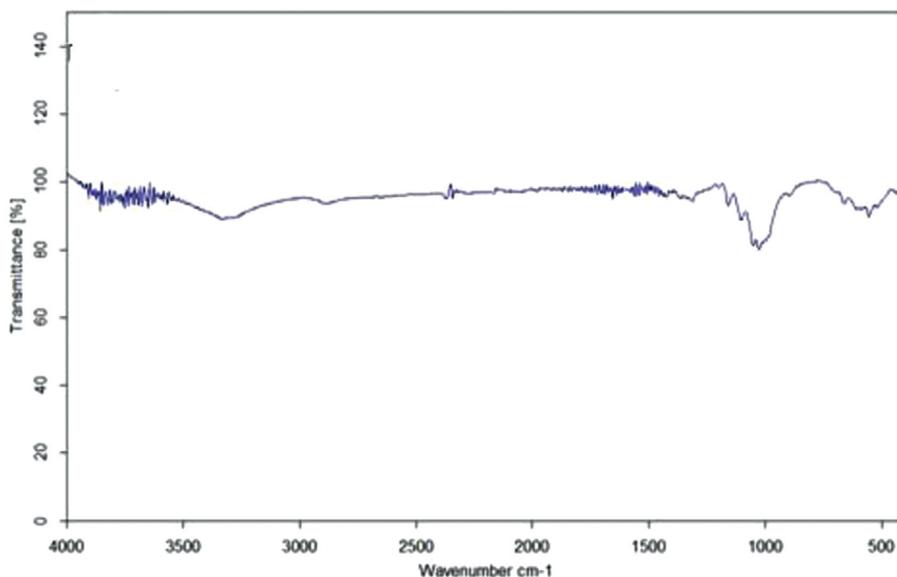


Figure 3: IR spectra of materials obtained by functionalized cellulose with crown ether DB18C6 and iron ions.

aromatic C = C bond stretching vibrations [32]. Bands characteristic of DB18C6 extractant appear in the range 1,550–500 cm^{-1} , the most intense ones are located at 1,000 cm^{-1} and 550 cm^{-1} [33].

Thus, dibenzo-18-crown-6-crown ether-specific vibration bands can be observed at around 1,000 cm^{-1} and 1,100 cm^{-1} , which might be attributed to calphatic-O-caromatic and calphatic-O-calphatic bonds, respectively [32]. Iron surface loading is evidenced by the presence of the band located at 1,037 cm^{-1} , which is characteristic for the Fe-OH bonds [30].

3.2. Sorption study results

3.2.1. The effect of contact time and temperature

The effect of contact time on the adsorption of As(V) on the functionalized material was studied as it was previously described. The adsorption capacity of the material, q ($\mu\text{g g}^{-1}$), was calculated using the following equation:

$$q = \frac{(C_0 - C_f) V}{m} \quad (1)$$

where C_0 is the initial concentration of As(V) from solution ($\mu\text{g L}^{-1}$), C_f is the residual As(V) concentration from solution ($\mu\text{g L}^{-1}$), V is the volume of solution (L) and m is the mass of the adsorbent material (g).

Obtained experimental data are shown in Fig. 4. For this study, contact time was varied between 0.5 and 8 h, and temperature is in the range of 298–318 K. From data depicted in Fig. 4, it can be observed that the adsorption capacity increases with the increase of contact time, and the equilibrium is reached after about 4 h for all temperatures used during the test.

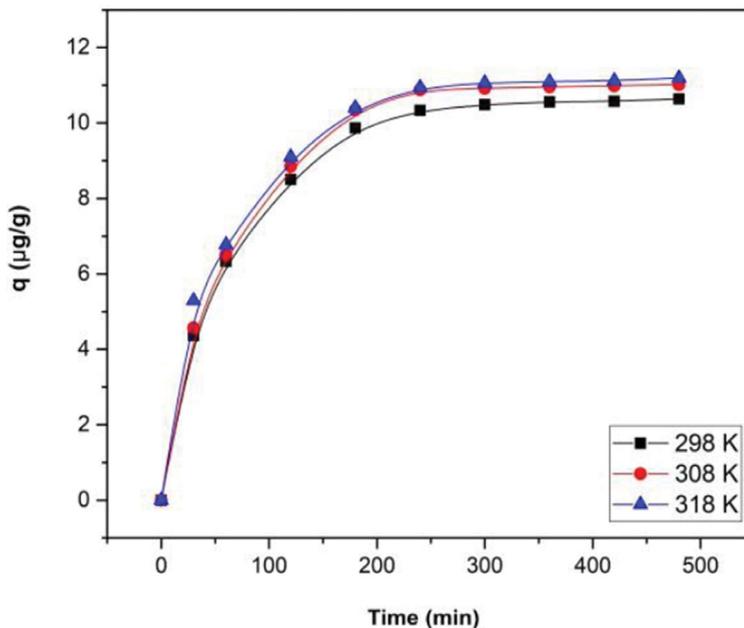


Figure 4: Effect of contact time on the adsorption capacity of the studied material at different temperature.

It might be observed that with increase in the temperature, the adsorption capacity of the material is not significantly affected.

3.2.2 Sorption kinetic studies

The kinetics of the As(V) adsorption process on material has been studied using two kinetic equations that could describe the process: pseudo-first-order kinetic equation proposed by Lagergren and the pseudo-second-order kinetic equation. The pseudo-first-order equation can be expressed as such:

$$\frac{dq_t}{dt} = k_1(q_e - q_t) \tag{2}$$

where q_e and q_t are the adsorbed amounts of As(V) per unit mass of material at equilibrium and time t , respectively, and k_1 is the rate constant for pseudo-first-order adsorption. The q_t at different times t can be determined by the following pseudo-first-order kinetic equation after integrating:

$$\ln(q_e - q_t) = \ln q_e - k_1 t \tag{3}$$

The pseudo-second-order kinetic model can be expressed by the following equation:

$$\frac{dq_t}{dt} = k_2(q_e - q_t)^2 \tag{4}$$

where k_2 is the rate constant for the pseudo-second-order adsorption.

By linearizing this equation, we obtained:

$$\frac{t}{q_t} = \frac{1}{k_2 q_e^2} + \frac{t}{q_e} \tag{5}$$

In Figs. 5 and 6, the pseudo-first-order and pseudo-second-order kinetic model plots obtained at the temperatures used during adsorption experiments (298, 308 and 318 K) as well the kinetic parameters k_1 and k_2 are depicted. For the modelling of experimental data, the linear shapes of the two models are used. The speed constant for the pseudo-first-order model was determined from the linear representation of $\ln(q_e - q_t)$ over time, and the velocity constant for the pseudo-second-order model was estimated from the linear representation of the t/q_t function of time.

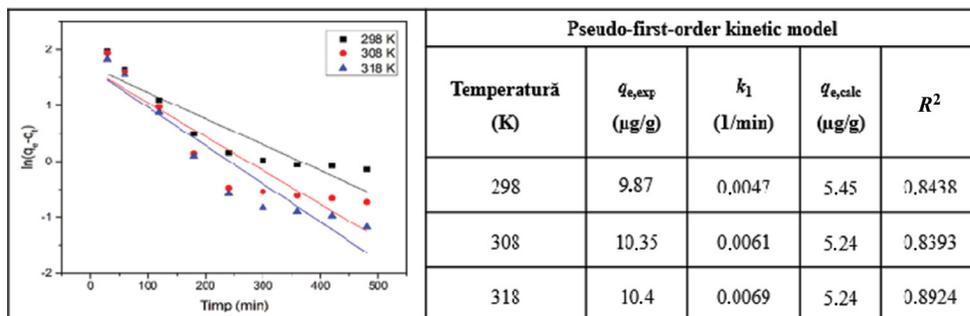


Figure 5: Pseudo-first-order kinetic model and experimental data.

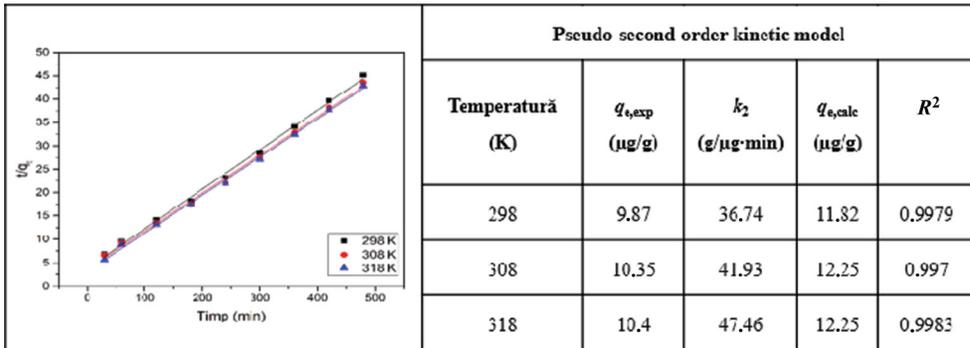


Figure 6: Pseudo-second-order kinetic model and experimental data.

Depending on the obtained values of constants and regression coefficients (R^2), the kinetic model that best describes the adsorption process can be established.

If the correlation coefficient R^2 is closer to 1, the adsorption process presents a better linearization for one of the two kinetic models presented. From the experimental data presented in the tables from Figs. 5 and 6, it is observed that pseudo-second-order model is better describing the arsenic adsorption onto the studied material, because the obtained correlation coefficient R^2 is in the range of 0.9970–0.9983 depending on the temperature and it is higher than that obtained for the pseudo-first-order model ($R^2 = 0.8438$ – 0.8924). The fact that adsorption kinetics was better described by the pseudo-second-order kinetic model suggested that the As(V) removal process corresponds to a chemisorption mechanism.

3.2.3 Sorption equilibrium studies

In order to establish the As(V) behaviour on the surface of the adsorbent material during the adsorption process, the experimental data obtained were modelled according to Langmuir, Freundlich and Sips isotherms (depicted in Fig. 7) which are generally used to describe the adsorption processes. Langmuir isotherm is applied for homogeneous surface adsorption

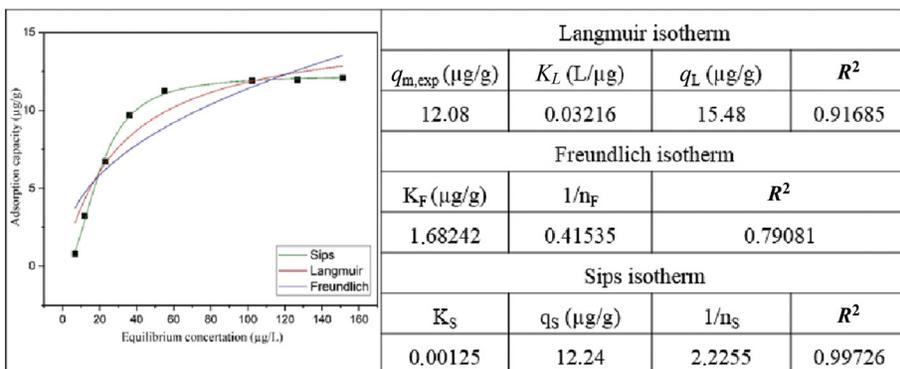


Figure 7: Adsorption isotherm of As(V) onto material and parameters of isotherm model.

[34–37]. Nonlinear expression of the Langmuir isothermal equation [39] can be expressed as follows:

$$q_e = \frac{q_{\max} \times K_L \times C_f}{1 + K_L \times C_f} \quad (6)$$

where q_e is the the maximum absorption capacity ($\mu\text{g L}^{-1}$), C_f is the equilibrium concentration or final concentration of As(V) in solution (mg/L), q_{\max} is the Langmuir maximum adsorption capacity ($\mu\text{g g}^{-1}$) and K_L is the Langmuir constant. The Freundlich isotherm can be applied for heterogeneous adsorption surfaces [35,38]. Nonlinear form of the Freundlich isotherm equation is:

$$q_e = K_F C_f^{1/n_f} \quad (7)$$

where q_e is the the maximum absorption capacity ($\mu\text{g g}^{-1}$), C_f is the equilibrium concentration or final concentration of As(V) in solution ($\mu\text{g g}^{-1}$), K_F and n_f are the characteristic constants that can be related to the relative adsorption capacity of the adsorbent and the intensity of adsorption. The Sips' isotherm is a combined form of the two previously presented models. Its nonlinear form [39] is the follows:

$$q_e = \frac{q_s K_S C_f^{1/n_s}}{1 + K_S C_f^{1/n_s}} \quad (8)$$

where q_s is the maximum absorption capacity ($\mu\text{g g}^{-1}$), K_S is the constant related to the adsorption capacity of the adsorbent and n_s is the heterogeneity factor.

The experimental data were fitted using eqns (6–8) and the studied isothermal parameters are shown in Fig. 7.

The correlation coefficient R^2 was calculated to establish which adsorption isotherm describes better the adsorption process of As(V). Figure 7 presents the parameters obtained for each isotherm used in the study. It can be observed that the highest value of the correlation coefficient R^2 was obtained when the experimental data were modelled using the Sips isotherm ($R^2=0.99726$). Thus, it can be concluded that the Sips model best described the adsorption process. It is also observed that the value of maximum adsorption capacity evaluated based on Sips isotherm model is $12.24 \mu\text{g As g}^{-1}$, much closer to the experimental value of $12.079 \mu\text{g As g}^{-1}$. The results indicated an increased adsorption capacity of the new material compared to chelating polymers loaded with Fe^{III} [40] or to silica used as solid support impregnated with crown ether and Fe^{III} used for the removal of As(V) from aqueous solutions [41]. Based on the results, it might be concluded that the adsorption process of As(V) on the obtained material is monolayer adsorption on a heterogeneous surface. The adsorption mechanism is controlled by chemisorption processes as a result of strong chelation between As(V) and OH^- groups or $\text{Fe}(\text{III})$ ions which are present on the surface of functionalized material. The value of the n_s coefficient with a value under 2 ($n_s < 2$) confirmed that the adsorption process is produced by the movement of As(V) ions from the aqueous phase to the material surface.

3.2.4. Thermodynamics of the adsorption

In order to investigate the spontaneity and thermodynamic properties of the adsorption process, the effect of the temperature on the adsorption process of As(V) was investigated by

varying the temperature in the range of 298–318K. Specific thermodynamic parameters were calculated as follows: Gibbs free energy (ΔG^0), free enthalpy (ΔH^0) and free entropy (ΔS^0) with the following relations:

$$\Delta G^0 = -RT \ln K_d \quad (9)$$

where:

$$K_d = \frac{C_{Ae}}{C_e} \quad (10)$$

and:

$$\log K_d = K_d = \frac{\Delta S^0}{2.3R} - \frac{\Delta H^0}{2.303RT} \quad (11)$$

where R is the universal gas constant ($8.314 \text{ J mol}^{-1} \text{ K}$), K_d is the distribution coefficient,

T is the absolute temperature (K), C_{Ae} is the equilibrium concentration of As(V) on adsorbent ($\mu\text{g L}^{-1}$) and C_e is the equilibrium concentration of As(V) in the solution ($\mu\text{g L}^{-1}$). Thermodynamic parameters in the case of As (V) adsorption on the obtained material were determined from the slope of the straight line and ordered at the origin of the linear representation of $\ln K_d$ as a function of $1/T$ and the values of ΔG^0 , ΔH^0 and ΔS^0 are shown in Fig. 8.

The correlation coefficient R^2 has a value equal to 0.9985. The positive value of enthalpy ΔH^0 demonstrates that the energy necessary for the adsorption process is the energy used to put in contact As(V) ions with the surface of the adsorbent material. The negative value obtained for free Gibbs energy suggests that the adsorption process of As(V) on the functionalized material is a spontaneous and natural process. These values become more negative with the increase of temperature, which can be attributed to effective growth of the contact surface of the adsorbent material and the As(V) ions. The positive value of entropy (ΔS^0) suggests that adsorption increases at the material/solution interface and the degree of particle clutter increases with increase in temperature, which can be attributed to some changes at the surface of the material. Thus, the adsorption of As(V) on the material surface is an endothermic and spontaneous process. Also, the value of activation energy has a value of $10.07 \text{ kJ mol}^{-1}$, suggesting that the adsorption of As(V) onto the studied material is a chemical adsorption.

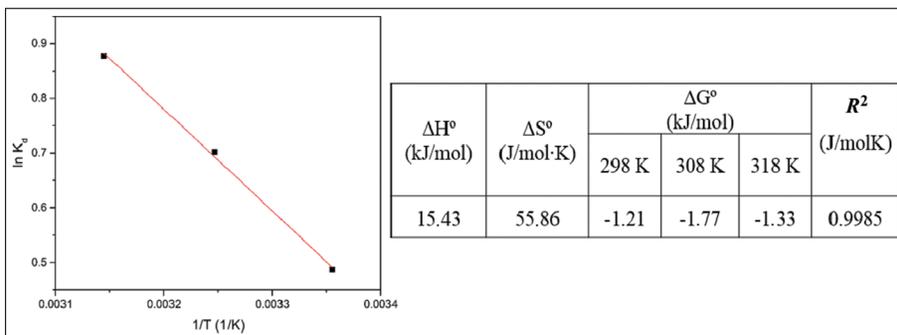


Figure 8: Arrhenius plot of As(V) adsorption onto adsorbent material.

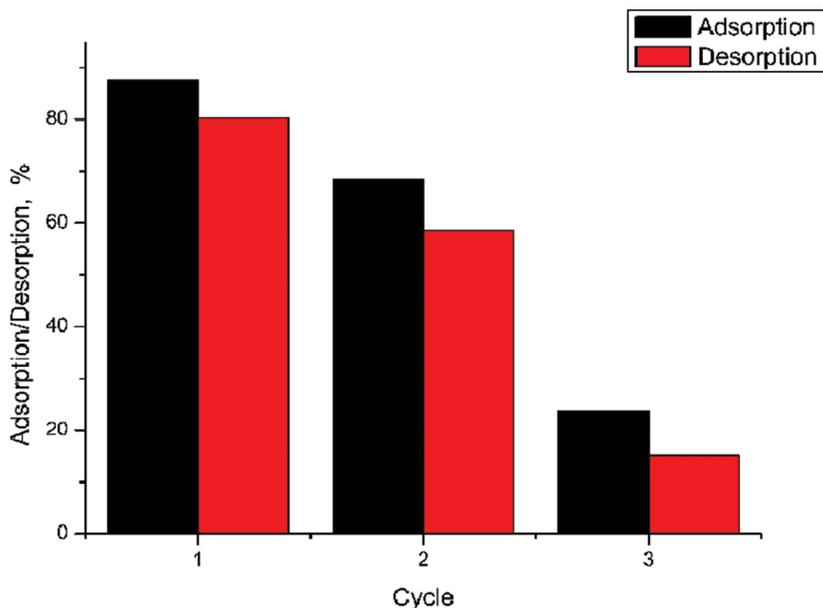


Figure 9: Adsorption–desorption cycles for As(V) uptake onto adsorbent material.

3.2.5 Adsorption–desorption experiments

Efficiency of the adsorbent material depends not only on his adsorption capacity but on his regenerative capacity and reuse as well. For this reason, in the present study, the possibility of reuse of the adsorbent material was investigated by establishing the adsorption/desorption cycle number (Fig. 9). Regeneration and reuse of the adsorbent material was performed by mixing 1 g material containing As(V) ions with 25 mL NaCl 5%. The mixture was shaken for 1 h at 200 rpm at room temperature, filtered and afterwards the material was dried at room temperature. This cycle was repeated until the As(V) ions were irreversibly fixed onto the surface of the adsorbent material, establishing in this way the maximum number of usage cycles.

It was noticed that the material efficiency decreases gradually, from an adsorption/desorption cycle to another, until the material is exhausted. The studies have shown that the adsorbent material can be reused for three adsorption/desorption cycles.

4 CONCLUSIONS

The results reported in this study have shown that the new adsorbent material obtained from cellulose as solid support functionalized with crown ether di-benzo-18 crown-6 (DB18C6) and loaded with Fe (III) ions had as a result an increase of As(V) removal efficiency from water. The new material was characterized by X-ray EDX technique and FTIR spectroscopy to highlight the changes occurred on the surface of cellulose particles after crown ether functionalization and iron ion loading.

The maximum adsorption capacity of the new adsorbent material was 12.08 $\mu\text{g As(V)/g}$ for an As(V) initial concentration of 60 mg/L. Kinetic studies showed that the optimum contact time between As(V) solution and functionalized material was of 240 min and the

optimum temperature was 298 K. Experimental data were modelled using the pseudo-first-order and the pseudo-second-order kinetic models. Based on the results, it was observed that the adsorption process is better described by the pseudo-second-order model.

Nonlinear regression analysis of the equilibrium data was performed using Langmuir, Freundlich and Sips isotherm models and the adsorption isotherm which best corresponds to the process is the Sips isotherm. The values of thermodynamic parameters were calculated and showed that the adsorption process was endothermic and spontaneous.

In order to study the efficiency of the new adsorbent material, adsorption–desorption experiments were carried out, and it was found that the material can be used for three cycles of adsorption/desorption

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